


CLICK ANYWHERE on THIS PAGE to RETURN to REMOVING STAINS on CONCRETE at InspectApedia.com

 An official website of the United States government



U.S. General Services Administration

Making Javelle Water

Procedure code:

371002S

Source:

Historic Concrete: Investigation & Repair- Pre-Conference Training Course, APT Annual Meeting, 1989.

Division:

Concrete

Section:

Concrete Cleaning

Last Modified:

08/11/2016

PART 1---GENERAL

1.01 SUMMARY

Javelle water is recommended for use in removal of the following stains:

- A. Beverage stains (see also "Removing Beverage Stains from Concrete").
- B. Fire, smoke, soot, pitch and wood tar stains (see also "Removing Fire, Smoke, Soot, Pitch And Wood Tar Stains From Concrete").
- C. Ink stains (see also "Poulticing Ink Stains From Concrete").
- D. Perspiration stains (see also "Poulticing Perspiration Stains from Concrete").
- E. Plywood or joint sealant stains (see also "Poulticing Plywood or Joint Sealant Stains from Concrete").
- F. Tobacco stains (see also "Removing Tobacco Stains from Concrete").

PART 2---PRODUCTS

2.01 MATERIALS

Note: Chemical products are sometimes sold under a common name. This usually means that the substance is not as pure as the same chemical sold under its chemical name. The grade of purity of common name substances, however, is usually adequate for stain removal work, and these products should be purchased when available, as they tend to be less expensive. Common names are indicated below by an asterisk (*).

A. Calcium Hypochlorite (CaCl₂O₂):

- 1. A white powder used especially as a bleaching agent and disinfectant.
- 2. Other chemical or common names include Chlorinated calcium oxide; Bleaching powder*; Calcium oxymuriate*; Chloride of lime*; Chlorinated lime*; Hypochlorite of lime*; Oxymuriate of lime*.

3. Potential Hazards: Caustic to flesh; flammable (when in contact with organic solvents); toxic from contact, inhalation & ingestion
4. Available from chemical supply house, dry cleaning supply distributor, drugstore or pharmaceutical supply distributor, janitorial supply distributor, swimming pool supply distributor, or water and sanitation supply distributor.

B. Sodium Carbonate (Na_2CO_3):

1. A sodium salt of carbonic acid used especially in making soaps and chemicals, in water softening, in cleaning and bleaching and in photography; A hygroscopic crystalline anhydrous strongly alkaline salt.
2. Other chemical or common names include Carbonate of soda*; Sal soda*; Soda*; Soda ash*; Washing soda*.
3. Potential Hazards: CAUSTIC TO FLESH.
4. Available from chemical supply house, grocery store or supermarket, hardware store, paint store or water and sanitation supply distributor.

C. Clean, potable water.

2.02 EQUIPMENT

- A. Enameled shallow pan.
- B. Rubber or plastic buckets or stoneware jar.

PART 3---EXECUTION

3.01 ERECTION, INSTALLATION, APPLICATION

- A. Dissolve 3 pounds of sodium carbonate crystals in 1 gallon of hot water in a rubber or plastic bucket.
- B. In a shallow enamel pan, place 12 ounces by weight of calcium hypochlorite. Add water slowly and mix to a paste while mashing the lumps.
- C. Pour the two solutions into another rubber or plastic bucket or into a stoneware jar and mix in enough water to make a total of 2 gallons of solution.
- D. Stir well, cover and allow the lime to settle.
- E. Carefully pour off the liquid for use, leaving the solids behind. The liquid can be siphoned off, but do not start the suction by mouth.

Last Reviewed: 2018-10-25