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Making a Solution for Removing Beverage, Soot, Tar and Other Stains from Concrete

Procedure code:

371001S

Source:

Historic Concrete: Investigation and Repair, APT Pre-Conference Training Course, APT Annual Meeting, 1989

Division:

Concrete

Section:

Concrete Cleaning

Last Modified:

12/26/2017

PART 1---GENERAL

1.01 SUMMARY

1. This standard includes guidance on making job-prepared hypochlorite for use in removing specific stains from concrete.

It is recommended for use in removal of the following stains:

1. Beverage Stains: See "Removing Beverage Stains from Concrete".
2. Fire, Smoke, Soot, Pitch and Wood Tar Stains: See "Removing Fire, Smoke, Soot, Pitch And Wood Tar Stains From Concrete".
3. Urine Stains: See "Poulticing Urine Stains From Concrete".
4. Wood Stains: See "Removing Wood Stains From Concrete".

PART 2---PRODUCTS

2.01 MATERIALS

NOTE: Chemical products are sometimes sold under a common name. This usually means that the substance is not as pure as the same chemical sold under its chemical name. The grade of purity of common name substances, however, is usually adequate for stain removal work, and these products should be purchased when available, as they tend to be less expensive. Common names are indicated below by an asterisk (*).

1. Calcium Hypochlorite (CaCl₂O₂):
 1. A white powder used especially as a bleaching agent and disinfectant.
 2. Other chemical or common names include Chlorinated calcium oxide; Bleaching powder*; Calcium oxy muriate*;

Chloride of lime*; Chlorinated lime*; Hypochlorite of lime*; Oxymuriate of lime*.

3. Potential Hazards: Caustic to flesh; flammable (when in contact with organic solvents); toxic from contact, inhalation & ingestion
 4. Available from chemical supply house, dry cleaning supply distributor, drugstore or pharmaceutical supply distributor, janitorial supply distributor, swimming pool supply distributor, or water and sanitation supply distributor.
2. Trisodium Phosphate:
1. Other chemical or common names include Tribasic sodium phosphate; Trisodium orthophosphate; Sodium orthophosphate; TSP*; Phosphate of soda*.
 2. Potential Hazards: Caustic to flesh.
 3. Available from chemical supply distributor, supermarket, grocery, or hardware store.
3. Clean, potable water

2.02 EQUIPMENT

1. Enameled shallow pan
2. Rubber or plastic buckets or stoneware jar

PART 3---EXECUTION

3.01 ERECTION, INSTALLATION, APPLICATION

1. Dissolve 2 pounds of sodium orthophosphate crystals in 1 gallon of hot water in a rubber or plastic bucket.
2. In a shallow enamel pan, place 12 ounces by weight of calcium hypochlorite. Add water slowly and mix to a paste while mashing the lumps.
3. Pour the two solutions into another rubber or plastic bucket or into a stoneware jar and mix in enough water to make a total of 2 gallons of solution.
4. Stir well, cover and allow the lime to settle.
5. Carefully pour off the liquid for use, leaving the solids behind. The liquid can be siphoned off, but do not start the suction by mouth.

Last Reviewed: 2018-10-25